

3. (9 pts) Two different compounds have the formula XeF_2Cl_2 .
- (4 pts) In the space below, draw Lewis structures for these two compounds

- (1 pt) Name the electron-pair geometry these compounds share:
- (1 pt) Name the molecular geometry these compounds share:
- (1 pt) Name the type of hybridization employed by the central atom in these compounds:
- (2 pts) Describe what property could be used to distinguish these two compounds experimentally.

4. (9 pts) Draw and name all of the isomers of C_4H_6 . Draw and name one isomer in each box below. All boxes need not be used, attach more paper if necessary.

Name:	Name:	Name:	Name:
Name:	Name:	Name:	Name:
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