

Quiz: Ch 1 & 2
AP Chemistry
Version M (33 pts)

Name:
Date: 8/26/05 **Period:**

Complete the following problems and record the answer with the correct number of significant figures. For scientific notation, “x 10^y” is used below. Show your work.

1. (3 pts) $805,715.475 / (56100 - 4285.5) =$

2. (3 pts) $268.5 + 490 + 3.80 \times 10^5 + 4.352 =$

3. (3 pts) Distinguish between the empirical formula and molecular formula of glucose, C₆H₁₂O₆.

Provide the name or chemical formula (as needed) for the following compounds. (2 pts each)

4. Cd(H₂PO₄)₂ _____
5. Hg₂Cr₂O₇ _____
6. S₈I₇ _____
7. HI _____
8. U(NO₃)₆ · H₂O _____
9. CO₃N₂ (“old school” name) _____
10. Selenous acid _____
11. Francium isothiocyanate _____
12. Plumbous manganate _____
13. Yttrium (III) sulfate _____
14. Tin (II) carbonite _____
15. Potassium monohydrogen arsenate _____

Quiz: Ch 1 & 2
AP Chemistry
Version N (33 pts)

Name:
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Complete the following problems and record the answer with the correct number of significant figures. For scientific notation, "x 10^y" is used below. Show your work.

1. (3 pts) $(500.2)(5.0000 \times 10^1) - 4652.50 =$

2. (3 pts) $3.80 \times 10^5 - 48.5 - 400 - 5623.352 =$

3. (3 pts) Distinguish between the atomic mass and mass number of potassium-39.

Provide the name or chemical formula (as needed) for the following compounds. (2 pts each)

4. $\text{Zn}(\text{HCO}_4)_2$ _____

5. HgO_2 _____

6. Se_4I _____

7. HF _____

8. $\text{AuBrO} \cdot 6 \text{H}_2\text{O}$ _____

9. Cu_2S ("old school" name) _____

10. Borous acid _____

11. Radium acetate _____

12. Stannic manganite _____

13. Platinum (II) cyanide _____

14. Xenon tetrafluoride _____

15. Ammonium bisulfate _____

Quiz: Ch 1 & 2
AP Chemistry
Version O (33 pts)

Name:
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Complete the following problems and record the answer with the correct number of significant figures. For scientific notation, “x 10^y” is used below. Show your work.

1. (3 pts) $(5000.2)(500. \times 10^1) - 4652.50 =$

2. (3 pts) $3.8 \times 10^5 - 4.5 - 4000 - 5623.2352 =$

3. (3 pts) Distinguish between the atomic mass and mass number of potassium-39.

Provide the name or chemical formula (as needed) for the following compounds. (2 pts each)

4. AgHCO_2 _____

5. Hg_2O_2 _____

6. Se_3I_7 _____

7. HBrO_4 _____

8. $\text{Au}(\text{IO})_3 \cdot 9 \text{H}_2\text{O}$ _____

9. CoS (“old school” name) _____

10. Phosphorous acid _____

11. Cesium oxalate _____

12. Nickel (II) chromite _____

13. Titanium (III) thiocyanate _____

14. Krypton hexachloride _____

15. Boron monoxide _____